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**Measuring**

The mole is the  
most common unit  
used to express

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the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles

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(or moles to  
volume) : at  $0^{\circ}\text{C}$   
and 1 atmosphere  
pressure (known  
as standard  
temperature and  
pressure or STP  
) , one mole of  
any gas occupies  
approximately  
22.4 liters.

**The Mole Unit -  
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Matter 1. pair

2. 5 3. dozen 4.

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1023 10. four

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mol Cu 12.4 23

4.1 mol CH<sub>4</sub> 6.02

10 molecules CH<sub>4</sub>

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A mole of a  
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substance is  
equal to the  
number of atoms

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in exactly 12 g  
of the carbon-12  
isotope.

Experiments  
established that  
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particles.

Avogadro's  
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mole equals

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6.022 \* 10<sup>23</sup>

items. The units  
of the mole are  
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atoms of each  
element in one  
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6.1.81 1024.

7.16.05 g/mol.

8.1.20 1024.

Section 10.4

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Formulas. 1. The  
percent

composition of a

compound is the

percent by mass

of each of the

elements in a

compound. 2.

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Divide the mass of each element in the sample by the mass of the sample. Then multiply each quotient by 100.

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the SI unit for  
measuring the  
amount of a

substance. The  
number of  
representative  
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, is called

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Chemists use the  
mole to count  
atoms,  
molecules, ions,  
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