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© OCR 2015 A Level in Chemistry A General Information . Molar gas volume = 24.0 dm³ mol⁻¹ at room temperature and pressure, RTP Avogadro constant, $N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$ Specific heat capacity of water, $c = 4.18 \text{ J g}^{-1} \text{ K}^{-1}$ Ionic product of water, $K_w = 1.00 \times 10^{-14} \text{ mol}^2 \text{ dm}^{-6}$ at 298 K 1 tonne = 10⁶ g Arrhenius equation: $k = A e^{-E_a/RT}$ or $\ln k = -E$

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of practical work with sample questions and answers help to improve exam technique in order to achieve higher grades. Written by experienced teacher Nora Henry, this Student Guide for practical Chemistry: - Help students easily identify what they need to know with a concise summary of required practical work examined in the A-level specifications. - Consolidate understanding of practical work, methodology, mathematical and other skills out of the laboratory with exam tips and knowledge check questions, with answers in the back of the book. - Provide plenty of opportunities for students to improve exam technique with sample answers, examiners tips and exam-style questions. - Offer support beyond the Student books with coverage of methodologies and generic practical skills not focused on in the textbooks.

Exam Board: OCR Level: A-Level Subject: Chemistry
First Teaching: September 2015 First Exam: Summer 2016
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